

#Jenny



Finally I get this ebook, thanks for all these I can get now!

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#Rio



Cool! I'am really happy

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#Markus Jensen



I did not think that this would work, my best friend showed me this website, and it does! I get my most wanted eBook

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#Hun Tsu



wtf this great ebook for free?!

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#Che Salsa



My friends are so mad that they do not know how I have all the high quality ebook which they do not!

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#Diego Butler



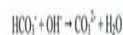
so many fake sites. this is the first one which worked! Many thanks

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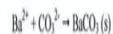
This procedure involves two titrations. First, total alkalinity (moles of bicarbonate + moles of carbonate) is measured by titrating the mixture with standard HCl to a methyl orange end point:



A separate aliquot of unknown is treated with excess standard NaOH to convert  $\text{HCO}_3^-$  to  $\text{CO}_3^{2-}$ :



Then all the carbonate is precipitated with  $\text{BaCl}_2$ :



The excess NaOH is immediately titrated with standard HCl to determine how much  $\text{HCO}_3^-$  was present. From the total alkalinity and moles of bicarbonate, you can calculate the original moles of carbonate present.

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